**COMMUNICATION**

CO₂ Electroreduction

**2D Metal Oxyhalide-Derived Catalysts for Efficient CO₂ Electroreduction**

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Electrochemical reduction of CO₂ to fuels and feedstocks, powered using renewable sources of electricity, contributes to closing the carbon cycle.[1–5] Specificity and productivity are measured via Faradaic efficiency (FE) and current density. Until now, the most productive electrocatalysts for formate—an attractive candidate as a liquid fuel and feedstock[6–9]—are derived from noble metals such as palladium and silver.[10–13]

Recent research efforts have focused on finding non-noble earth-abundant catalysts for formate production. Tin, bismuth, and lead have shown high stability and selectivity.[10,14–25] Despite the vast choice of catalysts, their productivity has generally not been sufficiently high.[22] They have suboptimal binding energies for CO₂ reduction intermediates and suffer from by-product poisoning at certain facets. The extensive surface reconstruction of metal catalysts under high current operating conditions makes it particularly challenging to program specific favorable facets, active sites, and oxidation states, during operation.[26–28]

Here, we report bismuth metal catalysts prepared via the in situ restructuring of 2D bismuth oxyhalides. The new catalysts exhibit enhanced specificity and productivity relative to previously reported bismuth catalysts.[29] They reach current densities above 200 mA cm⁻² at over 90% FE in the electroproduction of formate.

We achieved this by taking advantage of 2D layered precursor materials—bismuth oxyhalides. The favorable formation energy of this precatalyst leads to selective reconstruction of in situ electroreduction to form high-surface-area petal-structured electrodes. This results in the preferential exposure of more active Bi (110) facets, which we demonstrate with in situ grazing-incidence wide-angle X-ray scattering (GIWAXS) and X-ray absorption spectroscopy (XAS). This strategy increases the CO₂ reduction reaction specificity to over 90% while maintaining record-high current density (Table 1).

**Table 1.**

<table>
<thead>
<tr>
<th>Catalyst Type</th>
<th>FE (%)</th>
<th>Current Density (mA cm⁻²)</th>
</tr>
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<tbody>
<tr>
<td>BiOBr templated</td>
<td>90</td>
<td>200</td>
</tr>
<tr>
<td>Previous Bi</td>
<td>80</td>
<td>100</td>
</tr>
</tbody>
</table>

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In light of recent work reporting in situ formation of optimal CO₂RR catalysts via operando electro-redeposition,[26,27] we explored bismuth-based compounds as a materials platform with which we could deliberately control the morphology to promote efficient catalysis. The bismuth oxide sheets making up oxyhalides (BiOX), where X = I, Br, Cl, are separated by negatively charged halogens (Figure 1a).[30] We hypothesized that these materials would, upon electroreduction, generate highly active ultrathin layers of metallic bismuth as the halogen layer is dynamically removed.[29,31,32]

We prepared catalyst samples by coating the BiOB in dimethyl sulfoxide (DMSO) solution onto carbon paper electrodes and annealing in an inert atmosphere (see Experimental Section). The crystal structure of the resulting precatalyst was confirmed using X-ray diffraction, and matched well with the calculated peaks for this structure (Figure 1b). The precursor catalysts were then electroreduced in CO₂-saturated aqueous KHCO₃ solutions via 10 cycles of cyclic voltammograms. After this electroreduction preconditioning, the sample consisted of a combination of Bi and minor Bi₂O₃ species, which was then used as a catalyst without further modification. X-ray photoemission spectroscopy (XPS) reveals the stoichiometry reconfiguration after sample hydration toward a Br-rich BiOB, and the presence of residual Br in the reduced samples (Figure S1, Supporting Information). High-resolution transmission electron micrograph (HR-TEM) images reveal the stoichiometry reconfiguration after sample hydration toward a Br-rich BiOB, and the presence of residual Br in the reduced samples (Figure S1, Supporting Information). High-resolution transmission electron micrograph (HR-TEM) images reveal the stoichiometry reconfiguration after sample hydration toward a Br-rich BiOB, and the presence of residual Br in the reduced samples (Figure S1, Supporting Information).

Table 1. Performance comparison of various non-noble metals for CO₂ electroreduction to formate and best non-noble metal.

<table>
<thead>
<tr>
<th>Catalyst</th>
<th>Current density [mA cm⁻²]</th>
<th>Formate FE [%]</th>
<th>Citation</th>
</tr>
</thead>
<tbody>
<tr>
<td>BiOB templated</td>
<td>200</td>
<td>90+</td>
<td>This work</td>
</tr>
<tr>
<td>Metallic Bi</td>
<td>80</td>
<td>70</td>
<td>This work</td>
</tr>
<tr>
<td>Bi/ionic liquids</td>
<td>40</td>
<td>80</td>
<td>[19]</td>
</tr>
<tr>
<td>Bi/ionic liquids</td>
<td>12.5</td>
<td>90+</td>
<td>[16]</td>
</tr>
<tr>
<td>SnO nanoparticles</td>
<td>200</td>
<td>64</td>
<td>[37]</td>
</tr>
<tr>
<td>Sn</td>
<td>200</td>
<td>90</td>
<td>[15]</td>
</tr>
<tr>
<td>CO₃₂</td>
<td>10</td>
<td>90</td>
<td>[32]</td>
</tr>
<tr>
<td>Pb</td>
<td>62</td>
<td>95</td>
<td>[39]</td>
</tr>
<tr>
<td>Ag</td>
<td>300</td>
<td>60</td>
<td>[38]</td>
</tr>
</tbody>
</table>

Figure 1. Bismuth oxyhalide-templated catalysts. a) Atomic structure of BiOB and BiOB-templated Bi catalyst after electroreduction. b) X-ray diffraction patterns of BiOB precatalyst and BiOB-templated samples, which are obtained after hydration of BiOB/carbon paper. The peak at 26° is that of the carbon support. BiOB-templated catalysts are obtained by electroreducing BiOB in a CO₂-purged electrolyte. The crystal structure of BiOB-templated catalysts consists of a combination of Bi and Bi₂O₃ species. c,d) HRTEM before and after catalyst conditioning revealing the presence of different BiOB facets (hydrated, before reaction) and their selective reconstruction into Bi after operation. e,f) SEMs of samples after reaction reveal a 2D ultrathin petal-like arrangement after BiOB electroreduction.
To elucidate the structural evolution of the BiOX-templated catalyst at the atomic level, we carried out in operando GIWAXS measurements (Figure 2). BiOBr samples were placed in a custom-made reactor compatible with GIWAXS measurements and studied at different conditions (Figure 2a,b). At open-circuit conditions and in the absence of electrolyte, associated scattering peaks are seen in GIWAXS that we associate with BiOBr, Bi$_2$O$_3$, and Bi species (Figure 2c). After we add 1 M KHCO$_3$ electrolyte, the signal-to-noise ratio decreases and only the peak associated with BiOBr remains.

We then operated the catalyst under increasingly reducing conditions. At −1.5 V versus reversible hydrogen electrode (RHE), there is a significant catalyst reconfiguration; Bi (1T0) facets become the dominant arrangement, with a suppression of bismuth oxide species and a decreased contribution of BiOBr. This is in agreement with the Pourbaix phase diagram of BiOBr at this operating condition (Figure S2, Supporting Information), which predicts a configuration consisting of Bi (s) and Br- ions to be the most thermodynamically favorable.

To shed light on the electronic structure of the BiOBr-templated catalyst at operating conditions, we performed X-ray absorption near edge structure (XANES) and extended X-ray absorption fine structure (EXAFS) definition on BiOBr and Bi samples (Figure 2d). The absorption onset of the Bi L3 transition shifts toward higher energy for BiOBr samples, which indicates an increasing oxidation state of as-prepared BiOBr compared to control Bi. BiOBr-templated samples exhibit a reduction in the oxidation state after operation consistent with BiOBr electroreduction. We resolved the Fourier transformed $|\chi(R)|$ spectrum to evaluate the density of neighboring atoms as a function of radial distance (Figure 2e). A different distribution of Bi–O, Bi–Br, and Bi–Bi bonds is evident for Bi, BiOBr, and BiOBr-templated samples. Both oxybromide and oxybromide-derived samples exhibit a modified presence of closer-distance oxide neighbors and an apparent Bi–Bi shortening. BiOBr samples possess their strongest peak at $R = 1.8$ Å$^{-1}$, which we attribute to the presence of O and Br neighbors. BiOBr-templated samples show a reduction in the density of these states, in agreement with XPS measurements.
and instead exhibit two local maxima near 2.3 and 3.0 Å⁻¹. This is consistent with a Bi (110) configuration with an azimuthal compression, in contrast to bulk Bi control samples. We hypothesize that this favors the exposure of more active Bi facets as observed by in situ GIWAXS.

We then sought to evaluate the performance of the BiOBr-templated catalysts for CO₂ electroreduction. Testing of these catalysts in an aqueous H-cell set-up (0.1 M KHCO₃ electrolyte) revealed their higher activity compared to BiOₓ controls, exhibiting more than a twofold increase in current density, which led to 80 mA cm⁻² at −1.0 V versus RHE (Figure 3a). This approaches the CO₂ reduction current density limit achievable in the H-cell owing to CO₂ mass-transport limitations.[34]

To study the activity of the catalyst with respect to operating potential, we analyzed the selectivity of BiOBr catalysts toward formate at different reducing potentials (Figure 3b). At potentials more negative than −1 V, hydrogen evolution accounted for less than 3% of product and CO production was totally suppressed. The maximum HCOO⁻:H₂ ratio was observed in the range of −0.8 to −1.0 V versus RHE after iR correction, with the formate FE reaching values over 99%. Liquid product analysis was performed using NMR in water suppression mode (see Experimental Section). Bi control samples, on the other hand, exhibited selectivities to formate of about 70% in this potential range. BiOBr-templated catalysts exhibited remarkable stability, and retained their original performance during the course of an initial 65 h study at continuous operation (Figure 3c). This represents a twofold improvement in current density at near-unity FE compared to the best bismuth-based catalyst for CO₂ electroreduction into formate in traditional H-cell configurations. We also sought to switch the selectivity of BiBrO-templated catalysts toward CO and, when we operated it in an aqueous-free electrolyte, we achieved a similar productivity for this product (Figure S3, Supporting Information).

In view of the high current densities exhibited by BiOBr-templated catalyst, which reached CO₂ mass-transport limitations in an H-cell configuration, we developed catalysts that would work in a flow-cell system in which gas and liquid media are separated, overcoming mass transport in aqueous electrolytes and thus allowing a much higher CO₂RR current density and thus overall system productivity. To this end, we deposited BiOBr on top of a gas diffusion layer carbon electrode as a seed for the final BiOBr-templated catalyst. This allowed us to operate at much higher current densities in excess of 200 mA cm⁻² in a 2 M KHCO₃ electrolyte (Figure 3d and Figure S4, Supporting Information). The high selectivity toward formate remained largely unaffected, reaching a record value of 90% ± 5% even at these high current densities. Bi control
samples consistently exhibited lower current densities and formate selectivity around 70% ± 5%.

To provide further mechanistic insights into formate production on the BiOBr-templated catalyst, we performed density functional theory (DFT) calculations (Figure 4). Based on the operando GIWAXS experiments, we focused on the Bi (121) and Bi (110) facets to determine the effect of restructuring on the catalytic activity. From GIWAXS experiments, we found Bi (121) facet present in the precatalyst phase while Bi (110) emerged as the dominant facet during the operation. Only by measuring the surface facet structure under reaction conditions were we able to develop model metal slab models representative of the experimental catalyst. We sampled symmetric binding sites on the surface for both the HER (H*) and CO2RR to formate (HCOO*) intermediates (Figures S5 and S6, Supporting Information) and chose the lowest energy structures for further analyses.

We explored the reaction energy pathways for CO2 reduction to formate and the competing hydrogen evolution reaction (HER). Due to the near-unity experimental selectivity for formate, we did not focus on the CO production pathway. We found that the reaction energy barrier (∆G) was significantly higher for both HER (0.90 eV) and CO2RR (0.74 eV) on the Bi (121) facet. However, the Bi (110) facet displayed significantly lower reaction free energies for both HER (0.14 eV) and CO2RR (0.03 eV). Importantly, the reaction energy barrier for CO2RR to formate is the lowest of all reactions on the Bi (110) facet with the rate determining step being the second proton-coupled electron transfer to the bound HCOO* intermediate. Additionally, the stepped Bi (110) surface offers near-optimal Gibbs free binding energies for HCOO* (−0.01 eV) close to the thermodynamic minimum. The GIWAXS, electrocatalytic experiments, and DFT studies taken together toward the promotion of a highly active surface facet during CO2 reduction that favors formic acid production to the exclusion of competing products.

In conclusion, we demonstrate a new catalyst design strategy beginning with bismuth oxyhalides as a template from which we derived 2D bismuth-based materials, enabling the realization of catalysts that can sustain simultaneously high selectivity and activity for CO2 electroreduction. We employed in situ GIWAXS measurements to demonstrate that BiOBr-templated catalysts exhibit a preferential exposure of more highly active Bi (110) facets. We thereby achieve a high selectivity toward formate over 90% that remains up to current densities as high as 200 mA cm−2, demonstrating the potential of metal oxyhalide-templated catalysts for the efficient electroreduction of CO2.

**Experimental Section**

An extensive description of materials, sample fabrication and characterization, product analysis and computational studies, can be found in Supporting Information.

**Supporting Information**

Supporting Information is available from the Wiley Online Library or from the author.

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Conflict of Interest

The authors declare no conflict of interest.

Keywords

2D materials, catalysis, CO₂ electroreduction, formate, metal oxyhalides

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